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Chapter 8 Covalent Bonding. chemical bond. covalent bond. molecule. electron dot diagram. □□□the force that holds two

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atoms together. □□□the chemical bond that results from sharing valence electro.... □□a molecule is formed when two or more atoms bond covalently. □□□□used to show valence electrons of atoms.

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CHAPTER 8 SOLUTIONS MANUAL Covalent Bonding Covalent Bonding Solutions Manual Chemistry: Matter and Change • Chapter 8 121 Section 8.1 The Covalent Bond pages 240–247 Practice Problems page 244 Draw the Lewis structure for each molecule. 1. PH₃ H₂ H₂H— H H P respectively, for single, double, and triple P — — 2. H₂ S H H H — H S S — 3. HCl

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8.1 Molecular Compounds 8.2 The Nature of Covalent Bonding 8.3 Bonding Theories 8.4 Polar Bonds and Molecules

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1) Ionic bonds—transfer of electrons. (gained or lost; makes formula unit) 2) Covalent bonds—sharing of electrons. The resulting particle is called a “molecule”. Covalent Bonds. • The word covalent is a combination of the prefix co- (from Latin com, meaning “with” or “together”), and the verb. valere, meaning “to be strong”.

Chapter 8 “Covalent Bonding”

Chapter 8 Covalent Bonding and Molecular Structure 8-3 There are two types of repulsive forces between the two atoms. First, the nuclei repel because they are both positively charged. Second, the electrons repel because they are both negatively

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charged. The attractive forces between the two atoms result from the

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242 Chapter 8 • Covalent Bonding Single Covalent Bonds When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is

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often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

Chapter 8: Covalent Bonding

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2p* σ. 2p* O. 2 + (↑↓) (↑↓) (↑↓) (↑↓) (↑↓) (↑) () () Chapter 8: Covalent Bonding 297 As a result of how the electrons fill the MO diagram, the bond order is $(8 - 3)/2 = 2.5$, meaning that the bond between the N and the O is half way between a double bond and triple bond.

Chapter 8: Covalent Bonding Solutions for Chapter 8 ...

Chemistry Chapter 8- covalent bonding. STUDY. PLAY. covalent bond. the chemical bond resulting from sharing valence electrons. compounds. a substance formed when two or more elements are chemically bonded together. molecule. formed when two or more atoms bond covalently. single bond.

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8.2 The Nature of Covalent Bonding > 23 Copyright © Pearson Education, Inc., or its affiliates. All Rights Reserved. • Experimental evidence, however, indicates ...

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Chapter 8

8.6 Lewis Structures and Formal Charge • The electron surplus or deficit, relative to the free atom, that is assigned to an atom in a Lewis structure. Formal charges are not “real” charges. H: orig. valence e = 1 non-bonding e = 0 1/2 bonding e = 1 formal charge = 0 O: orig. valence e = 6 non-bonding e = 4

Chapter 8 Chemical Bonding I: Basic Concepts

8.4: Bond Polarity and Electronegativity Bond polarity - describe the sharing of electrons between atoms Nonpolar covalent bond - one in which the electrons are shared equally between two atoms (Cl₂, H₂, N₂, etc.) Polar covalent bond - one of the atoms exerts a greater attraction for the bonding electrons than the other