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6. IE of Kr > IE of Br because they are in the same valence shell and Kr has the higher core charge (+8 vs. +7). IE of Rb is the lowest because core charge is +1 and its valence shell ($n = 5$) is larger than the valence shell ($n = 4$) of Kr and Br.

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Isotope Atomic Number Z Mass Number A Number of Electrons 27 59 25 14 N 3 7 3 3 6 3 58 Zn²⁺ 19 F⁻ 10. Using grammatically correct English sentences, describe what the isotopes of an element have in common and how they are different. 11. J. N. Spencer, G. M. Bodner, and L. H. Rickard, Chemistry: Structure & Dynamics,

The Nuclear Atom

Show your work. ChemActivity 2 Atomic Number and Atomic Mass 11 10. Individually, use your

answer to CTQ 9b to calculate the mass (in grams) of 6,022 x 10²³ (randomly selected) chlorine atoms. Once all group members have completed the calculation, compare your answers and come to consensus. 11.

Solved: ChemActivity 2 Atomic Number And Atomic Model 2: I ...

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ChemActivity 6 Atomic Size 372. Why does the core charge increase as one moves from left to right across a period in the periodic table—for example, from boron to carbon to oxygen?3. What trend in atomic radius is observed as one moves from left to right across a period? Why?4.

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