

### Section 71 Ionic And Metallic Bonding Answers

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#### Section 71 Ionic And Metallic

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Chapter 7 Ionic and Metallic Bonding59 SECTION 7.1 IONS (pages 187–193) This section explains how to use the periodic table to infer the number of valence electrons in an atom and draw its electron dot structure. It also describes the formation of cations from metals and anions from nonmetals. Valence Electrons (pages 187–188) 1.

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## **SECTION 7.1 IONS (pages 187-193)**

Chapter 7 "Ionic and Metallic Bonding" Pre-AP Chemistry Charles Page High School Stephen L. Cotton Section 7.1 - Ions OBJECTIVES: - Determine the number of valence electrons in an atom of a representative element.

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## **Section 7 Ionic Metallic Bonding Answers [EBOOK]**

Ionic bonding is a type of chemical bond that occurs between two oppositely charged ions while metallic bonding is the type of chemical bond that occurs in a metal lattice. Hence, the key difference between ionic bonding and metallic bonding is that the ionic bonding takes place between positive and negative ions whereas the metallic bonding takes place between positive ions and electrons.

## **Difference Between Ionic Bonding and Metallic Bonding ...**

Essential Understanding Ionic compounds are the result of ionic bonds forming between oppositely charged ions. Lesson Summary Formation of Ionic Compounds An ionic compound is made up of anions and cations and has an overall charge of 0. The electrostatic attraction between an anion and a cation is an ionic bond.

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## **BONDING AND INTERACTIONS**

Metallic bond: Metallic bonds are forces between negatively charged freely moving electrons and positively charged metal ions. Bond Energy. Ionic Bonds: Bond Energy is higher than metallic bonds. Covalent Bonds: Bond Energy is higher than metallic bonds. Metallic Bonds: Bond Energy is lower than other primary bonds. Formation

## **Difference Between Ionic Covalent and Metallic Bonds ...**

Section 7.3 – Bonding in Metals. The valence electrons of metal atoms can be modeled as a sea of electrons. Metallic bonds consist of the attraction of the free-floating valence electrons for the positively charged metal ions. Metals are good conductors and malleable because of their mobile electrons.

## **Chapter 7 - Ionic and Metallic Bonding**

Section 8.2 What is an ionic bond? In your textbook, read about forming ionic bonds and the characteristics of ionic compounds. Circle the letter of the choice that best completes the statement or answers the question. 1. An ionic bond is a. attraction of an atom for its electrons. b. attraction of atoms for electrons they share.